

SSC PUBLIC EXAMS - TELANGANA STATE

MODEL PAPER-III

Sub : Physical Science

CLASS : X

PAPER-I

Time : 2 hours 45 min

Max. Marks : 40

Instructions :

1. In the time duration of 2 hours 45 minutes, 15 minutes of time is allotted to read and understand the question paper.
2. Answer the questions under PART-A on separate answer book
3. Write the answers to the questions under PART-B on the question paper itself and attach it to the answer book of PART-A

Part-A

Time : 2 hours 15 min

Max. Marks : 40

Instructions :

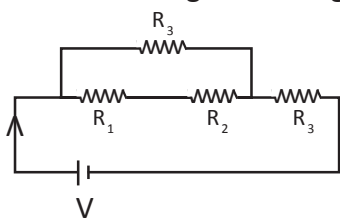
1. PART-A comprises of three sections I, II, III
2. All the questions are compulsory
3. There is no overall choice. However, there is internal choice to the questions under section III

Section-I

Notes :

7x1=7

1. Answer all the questions
 2. Each question carries 1 mark
 3. Answer each question in 1 or 2 sentences
1. How can you differentiate between strong acid and strong base?
 2. Which principles refer for He electronic configuration orientation?
 3. This circuit has given wrong away, correct it finally we want series connection.



4. Arrange the low reactivity metals like ascending order Pt, K, Zn, Ag, Hg, Fe, Au
5. How the Glycerol formed?
6. Explain about the formation of Saponification reaction?
7. Write the reasons for dry HCl is Acid or base?

Section-II

Notes :

6x2=12

1. Answer all the questions
 2. Each question carries 2 mark
 3. Answer each question in 4-5 sentences
8. Draw the ray diagrams for image formation, when one object is keeping in different places in front of convex lens?
 9. How the four quantum number are helpful to "Ca" Element in the position of outmost last electron?
 10. What type of chemical reaction takes place in breath analyser, that is helped for police to detect suspect drunken driver?

11. Draw the Electrical circuit with help of these values, $R_1=4\Omega$, $R_2=12\Omega$ are parallel connection and $R_3=7\Omega$ in series connection these are connect to 12V Battery?
12. Write predictions for all metals will be getting corrosion?
13. Write the bond formation of $MgCl_2$ using cation formation and Anion formation?

Section-III

Notes :

6x2=12

1. Answer all the questions
2. Each question comes 4 mark
3. There is an internal choice for each question. only re option from each question is to be attempted
4. Answer each question in 8-10 sentences

14. Write about kirchhoff's laws? like junction law and loop law?

(OR)

How can you use the valence bond theory for formation of $BeCl_2$.

15. If Modi wants to protect his eye, give the suggestions about each and every part of eye parts. briefly explain it?

(OR)

Draw the ray diagrams with help of the following Table.

Position of Candle	Positon of Image	Enlarged Diminished	Inverted Erect
Beteen "F" and "C"	Beyond "C"	Enlarged	Inverte
Beyand "C"	Beteen "F" and "C"	Diminished	Inverte
Between mirror & "F"	Behind the mirror	Enlarged	Erect

16. Write the procedure for dilute HCl is react with Zinc granuales?

(OR)

Write the procedure for D.C Generator?

17. Determine the Magnetic force a Current carrying wire, which is placed along a uniform magnetic field?

(OR)

Write the steps to balance the chemical Equation ex; $Fe_2O_3 + Al \rightarrow Fe + Al_2O_3$.

Part-B

Time : 30 min

Max. Marks : 10

Instructions :

1. Answer all the questions
2. Each question carries 1/2 mark
3. Answers are to be written in the question paper only.
4. Marks will not be awarded in any case of overwriting rewriting or erased answers.
5. Write capital letter (A, B,C,D) Showing the current answer for the following questions in the brackets against them.

-
18. Speed of light (C)=..... ()
a) $3 \times 10^8 \text{ ms}^{-1}$ b) $3 \times 10^{12} \text{ ms}^{-1}$ c) $3 \times 10^{10} \text{ ms}^{-1}$ d) $3 \times 10^9 \text{ ms}^{-1}$
 19. Why does the sky appears as white in hot day. ()
a) Scatlering of light b) Dispersion c) Rafraction d) Reflection
 20. Lens maker's formula ()
a) $\frac{1}{f} = (n-1)\left(\frac{1}{R_1} - \frac{1}{R_2}\right)$ b) $\frac{1}{f} = (1-n)\left(\frac{1}{R_1} - \frac{1}{R_2}\right)$ c) $\frac{1}{f} = (n-1)\left(\frac{1}{R_2} - \frac{1}{R_1}\right)$ d) $\frac{1}{f} = (1+n)\left(\frac{1}{R_1} - \frac{1}{R_2}\right)$
 21. One solution turns Red to Blue its P^H is likely to be. ()
a) 1 b) 3 c) 7 d) 14
 22. Bleaching powder formula ()
a) CaOCl_2 b) Ca(OH)_2 c) CaO d) Na_2CO_3
 23. Avagrado Number = ()
a) $6.02 \times 10^{23} \text{ NA}$ b) $60.2 \times 10^{23} \text{ NA}$ c) $6.20 \times 10^{23} \text{ NA}$ d) $6.02 \times 10^{32} \text{ NA}$
 24. The radius of curvaure will be twice to this distnace ()
a) $2f$ b) $f/2$ c) $f/4$ d) F
 25. Example for amorphous forms ()
a) Lamp black b) Buck minster fullerence c) Diamond d) Graphite
 26. The impurity present in the ore is called as..... ()
a) Mineral b) flux c) Gangue d) Slag
 27. This law based on the conservation of charge ()
a) Junction law b) Looplaw c) Octet rule d) Law of Triads