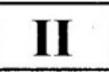
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Total	No.	of	Questions – 21				
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Part - III CHEMISTRY, Paper-II (English Version)

Time: 3 Hours | Max. Marks: 60

Note: Read the following instructions carefully:

- Answer all questions of Section 'A'. Answer any six questions in Section 'B' and any two questions in Section 'C'.
- (2) In Section 'A', questions from Sr. Nos. 1 to 10 are of "Very short answer type". Each question carries two marks. Every answer may be limited to two or three sentences. Answer all these questions at one place in the same order.
- (3) In Section 'B', questions from Sr. Nos. 11 to 18 are of "Short answer type". Each question carries four marks. Every answer may be limited to 75 words.
- (4) In Section 'C', questions from Sr. Nos. 19 to 21 are of "Long answer type". Each question carries eight marks. Every answer may be limited to 300 words.
- (5) Draw labelled diagrams, wherever necessary for questions in Section 'B' and 'C'.

SECTION - A

 $10 \times 2 = 20$

Note: Answer all the questions:

- Define Osmotic Pressure.
- What are antibiotics? Give example.
- 3. Write the difference between a soap and a synthetic detergent.
- 4. State Faraday's first law of electrolysis.
- Give the composition of the following alloys:
 - (a) Brass
- (b) German Silver
- 6. List out any two uses of Neon.
- Write the structure of XeF₄.
- 8. Why Zn²⁺ is diamagnetic whereas Cr³⁺ is paramagnetic?
- What is biodegradable polymer? Give one example.
- 10. Write the names of the monomers of the following polymers:
 - (a) Bakelite
- (b) Nylon 6,6

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Note: Answer any six questions:

- 11. What is doping? Explain n-type and p-type semi-conductors.
- 12. Explain the purification of Sulphide ore by froth floatation method.
- 13. State Raoult's law. The vapour pressure of pure Benzene at a certain temperature is 0.850 bar. A non-volatile, non-electrolyte solid weighing 0.5 g when added to 39.0 g of benzene (molar mass 78 g mol-1). Vapour pressure of the solution, then is 0.845 bar. What is the molar mass of the solid substance?
- 14. Describe the purification of colloidal solutions by Dialysis method with a neat diagram.
- 15. Write the equations for reactions of chlorine with the following
 - (a) Cold and dilute NaOH
- (b) Excess NH₃

(c) Ca(OH),

- (d) Na₂S₂O₃
- Explain Werner's theory of Co-ordination compounds.
- 17. What are Hormones? Give one example for each of the following:
 - (a) Steroid hormones
 - (b) Polypeptide hormones
 - (c) Amino acid derivatives
- 18. Explain the mechanism of S_N2 reaction with one example.

SECTION - C

 $2 \times 8 = 16$

Note: Answer any two of the following questions:

- 19. (a) What are Galvanic cells? Explain the working of Galvanic cells with one example.
 - (b). Write the differences between order and molecularity of a reaction.
- 20. (a) How is ammonia manufactured by Haber's process? Explain.
 - (b) How does ozone react with the following:
 - (i) PbS
- (ii) KI
- (iii) C2H4
- (iv) NO

- 21. Explain the following reactions with suitable example:
 - (a) Reimer-Tiemann reaction
 - (b) Cannizaro reaction
 - (c) Aldol condensation
 - (d) Sandmeyer reaction