

10. Chemical Bonding

1. Ionic compounds dissolved in ____ solvents.
2. Example for polar solvents is ____.
3. The shape of BeCl_2 is ____.
4. Electro positivity is also called as ____.
5. Valence bond theory was proposed by ____.
6. The shape of NH_3 is ____.
7. 1 nanometer = ____Meter
8. Examples of Triple bond molecules ____.
9. General electronic configuration of Noble gas ____.
10. The noble element which is exception of octet rule ____.
11. Which of the following elements is electronegative? ()
a) Sodium b) Oxygen c) Magnesium d) Calcium
12. An element ${}_{11}^{23}\text{X}$ forms an ionic compound with another element Y. then the charge on the ion formed by X- is ()
a) +1 b) +2 c) -1 d) -2
13. An element 'A' forms chloride ACl_4 . The no. of electrons in the valence shell of A. ()
a) 1 b) 2 c) 3 d) 4
14. Bond angle in methane is ()
a) 170° b) 104.5 c) 109.28° d) 180°
15. Shape of ammonia molecule ()
a) Linear b) Angular c) Pyramidal d) Tetrahedral

Answers

1) Polar

2) H₂O

3) Linear

4) Metallic Character

5) Linus Pauling

6) Pyramidal

7) 10⁻⁹

8) N₂, C₂H₂

9) ns²np⁶

10) Helium

11) b

12) a

13) d

14) c

15) c