Periodic Classification of Elements

1. An element with Z= 20 will be placed in which period of the periodic table? (CBSE2001)

1) 4 2) 3 3) 2 4) 1

2. If the atomic number of an element is 33, it will be placed in the periodic table in the (MLNR)

1) First group 2) Third group

3) Fifth group 4) Seventh group

3. The electronic configuration of chromium in its ground state I (E-1997)

1) [Ar] $4d^4\ 4s^2$ 2) [Ar] $4d^5\ 4s^1$

3) [Ar] $3d^4 3s^2$ 4) [Ar] $3d^5\ 4s^1$

4. As per the modern periodic law, the physical and chemical properties of elements are periodic functions of their (E-1998)

1) atomic number 2) electronic configuration

3) atomic weight 4) atomic size

5. The number of periods present in the long form of the periodic table (E-1999)

1) 6 2) 7 3) 8 4) 18

6. The electronic configuration of group III elements is (E-2000)

1) $ns^2\ np^3$ 2) ns^2np^5 3) ns^2np^1 4) ns^2np^2

7. Which of the following pairs of ions have the same electronic configuration (E-2002)

1) Cr^{+3}, Fe^{+3} 2) Fe^{+3}, Mn^{+2} 3) Fe^{+3}, Co^{+3} 4) Sc^{+3}, Cr^{+3}

8. Which one of the following is correct about stability of the given ions (E-2003)

1) $Pb^{2+} > Pb^{4+}$ 2) $Pb^{4+} > Pb^{2+}$

3) $Si^{2+} > Si^{4+}$ 4) $Sn^{4+} > Sn^{2+}$

9. Which of the following is not an actinide? (DPMT2005)

1) Curium 2) Californium 3) Uranium 4) terbium

10. Europium is: (DPMT2005)

1) s-block (2) p-block 3)d-block 4)f-block

11. Element with atomic number 56 belongs to which block? (AFMC2002, 09)

1) s 2)p 3)d 4)f

12. which of the following sets is of coinage metals? (DPMT2003)

1) Cu, Ag, Au 2) Zn, Cd, Hg 3) Au, Ag, Zn 4) Li, Na, K

13. An element X belongs to fourth period and fifteenth group of the periodic table. Which of the following is true regarding the outer electronic configuration of X? It has: (PMT2008)

1) Partially filled d-orbital's and completely filled p-orbital

2) completely filled s-orbital and completely filled p-orbital's

3) completely filled s-orbital's and half filled p-orbital's

4) half-filled d-orbital's and completely filled s-orbital

KEY

1)2	**2) 3**	**3) 4**	**4) 2**	**5)2**	**6)3**	**7) 2**	**8)1**
9) 4	**10)4**	**11) 1**	**12) 1**	**13) 4**			