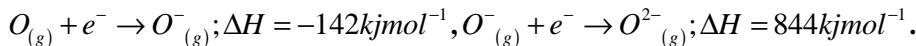


ELECTRON AFFINITY

1. The electron affinity values (KJmol^{-1}) of three halogens X,Y and Z are respectively -349, -333 and -325. Then X,Y and Z respectively are (M2003)

1) F_2 , Cl_2 and Br_2 2) Cl_2 , F_2 and Br_2 3) Cl_2 , Br_2 and F_2 4) Br_2 , Cl_2 and F_2

2. The formation of the oxide ion $\text{O}_{(g)}^{2-}$ requires first an exothermic and then an endothermic step as shown below: (AIEEE-2004)



This is because:

- 1) O^- ion has comparatively larger size than oxygen atom
- 2) Oxygen has high electron affinity
- 3) O^- ion will tend to resist the addition of another electron.
- 4) Oxygen is more electronegative

3. An element "X" has IP = 1681 kJ/mole and EA = -333 kJ/mole then its electro negativity is (AIIMS)

$$1) \frac{1681+333}{544} \quad 2) \frac{1681+333}{129} \quad 3) \frac{1681+333}{5.6} \quad 4) \frac{1681+333}{2.8}$$

4. If the E.N value of Nitrogen in the Pauling scale is 3.0, then the value in mulliken scale will be (BHU2003)

- 1) 11.2 2) 22.4 3) 8.4 4) 3.0

Hint-Mulliken EN value=2.8 X Pauling EN values

5. For Electron affinity of halogens, which of the following is correct? (AIIMS2004)

- 1) Br>F 2) F>Cl 3) F>I 4) Br>Cl

6. which of the following has highest Electron affinity? (BHU2005)

- 1) N 2) Cl 3) O 4) F

7. The correct order of electronegativity of the elements of halogen family is (PMT2008)

- 1) I > Br > F > Cl 2) Br > I > Cl > F 3) F > Cl > Br > I 4) Cl > Br > I > F

8. The correct order of decreasing electronegativity of the elements,i-Be,ii-O,iii-N and iv-Mg (PMT2011)

- 1) ii> iii >i > iv 2) iii>iv> ii >i 3) i > ii > iii > iv 4) i > ii >iv>iii

9. Increasing order of the electronegativity of elements in the following (AFMC2010)

- 1) C, N, Si, P 2) N, Si, C, P 3) Si, P, C, N 4) P, Si, N, C

10. Which of the following is 2nd most Electro negative element? (PMT2007)

- 1) S 2) Cl 3) O 4) F

KEY

- 1) 2 2) 3 3) 1 4) 3 5) 3 6) 2 7) 3 8) 1 9) 3 10) 3