CEEP MODEL PAPER 2012

TIME : 30 Min MAX MARKS : 30

Section – III
CHEMISTRY

91. Value of Plank's constant is
1) 6.625×10^{-27} erg/sec 2) 6.625×10^{-27} erg – sec
3) 6.625×10^{-34} erg/sec 4) 6.625×10^{-34} J/sec

92. The region where the probability of finding electron becomes zero is
1) Orbital 2) Orbit 3) Nodal plane 4) none

93. Factors affecting the Ionization potential
1) Nuclear attraction 2) atomic size 3) charge of the ion 4) all the above

94. What are the correct *n. l. m. s* values for the fourth electron of Beryllium?
1) 1,0,0, +1/2 2) 1,1,1,+1/2 3) 2,0,0, -1/2 4) 2,1,0, +1/2

95. The radius of the second stationary orbit is
1) $r = h/mv\pi$ 2) $r = 2h/mv\pi$ 3) $r = h/mv2\pi$ 4) none

96. Which of the following is a covalent compound?
1) $MgCl_2$ 2) LiF 3) N2 4) NaCl

97. In [NH_3] $\rightarrow$ BF_3 molecule, the donor is
1) H^+ ion 2) Boron 3) Nitrogen 4) Fluorine

98. Eka – Aluminium was discovered as
1) Gallium 2) Scandium 3) Germanium 4) Silicon

99. The general electronic configuration of d – block elements is
1) $ns^2 (n-1) d^{1-10}$ 2) ns^2 n d^{1-10}
3) $ns^2 (n-2) d^{1-10}$ 4) ns^2 nf^{1-14} d^{1-10}

100. Which of the following elements shows the strongest reducing property?
1) Li 2) Na 3) B 4) Be

101. The gas released by the action of MgO with Cl_2 gas at 1000^0C
1) SO_2 2) H2 3) CO 4) CO2

102. Which of the following reaction takes place at Cathode in electrolysis of $MgCl_2$?
1) $Mg^{2+} + 2e^- \rightarrow Mg$ 2) $Mg^{2+} \rightarrow Mg + 2e^-$
3) $2Cl^- + 2e^- \rightarrow Cl_2$ 4) $2Cl^- \rightarrow Cl_2 + 2e^-$

103. Find the volume of 2M of H_2SO_4 required to neutralize the solution of 1.5 litres with molarity 5.

1) 1.5 litres 2) 2.3 litres 3) 1.875 litres 4) 1.2 litres

104. If gms of urea is dissolved in 200 ml of its solution, Molarity is (At. Wt of urea is 60) 1) 0.2M 2) 0.5M 3) 0.05M 4) 0.1M

105. 25 gms of NaOH is prepared as 250 ml of solution, then its molarity is
1) 0.25M 2) 0.025M 3) 2.5M 4) 0.625M

106. If 500 ml 0.5 HCl solutions is changed to 0.1M strength, find out the volume of water to be added
1) 2000 ml 2) 2500 ml 3) 1000 ml 4) 1500 ml

107. If water is added to 100ml of 0.3 M HCl solution, then molarity attains as 0.1M. Then the volume of water added is
1) 100 ml 2) 200ml 3) 300 ml 4) 50ml

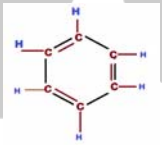
108. The volatility order of HCl, H_2So_4, CH_3COOH is
1) Hcl > CH_3COOH > H_2SO_4 2) CH_3COOH > H_2SO_4 > HCL
3) H_2SO_4 > HCL > CH_3COOH 4) H_2SO_4 < HCL < CH_3COOH

109. If the PH of a solution is changed from 7 to 14, then the strength of the Base
1) Increased 2) Decreased 3) No change 4) Can't say

110. The bond angle in diamond is
1) 120^0 2) $109^0\ 28'$ 3) 180^0 4) 107^0

111. Among the following the hexyl group is
1) C_6H_{12} 2) C_4H_6 3) C_6H_{13} 4) C_6H_{10}

112. The general formula of Marsh gas is
1) C_2H_6 2) C_3H_8 3) CH_{12} 4) CH_4

113.  The picture is
1) Hexyne 2) Hexene 3) Benzene 4) Hexane

114. The following gas is passed through sugar juice to remove the excess lime
1) SO_2 2) CO23) SO_3 4) CO

115. CO – NH bond is called
1) Ethylene bond 2) Acetylene bond
3) Peptide bond 4) Coordinate covalent bond

116. Protein present in nails
1) Haemoglobine 2) Keratin 3) Collagen 4) Insulin

117. The pressure required for hydrolysis of oil/fat is
1) 4 Pascal 2) 4.1 Mega Pascal 3) 1 Pascal 4) none

118. The protein which regulates blood sugar level
1) Insulin 2) Fibroid 3) Myosin 4) Keratin

119. The word 'Ceramics' is derived from
1) Feldspar 2) Atomus 3) Keramos 4) None

120. Thermo elastic plastic is
1) Cellulose Acetate 2) Phenol Formaldehyde
3) Urea – formaldehyde 4) Ipaxy resins

SECTION – III
CHEMISTRY

91)	2	101)	3	111)	3
92)	3	102)	1	112)	4
93)	4	103)	3	113)	3
94)	3	104)	2	114)	1
95)	1	105)	3	115)	3
96)	4	106)	1	116)	2
97)	3	107)	2	117)	2
98)	1	108)	4	118)	1
99)	1	109)	1	119)	3
100)	2	110)	2	120)	1